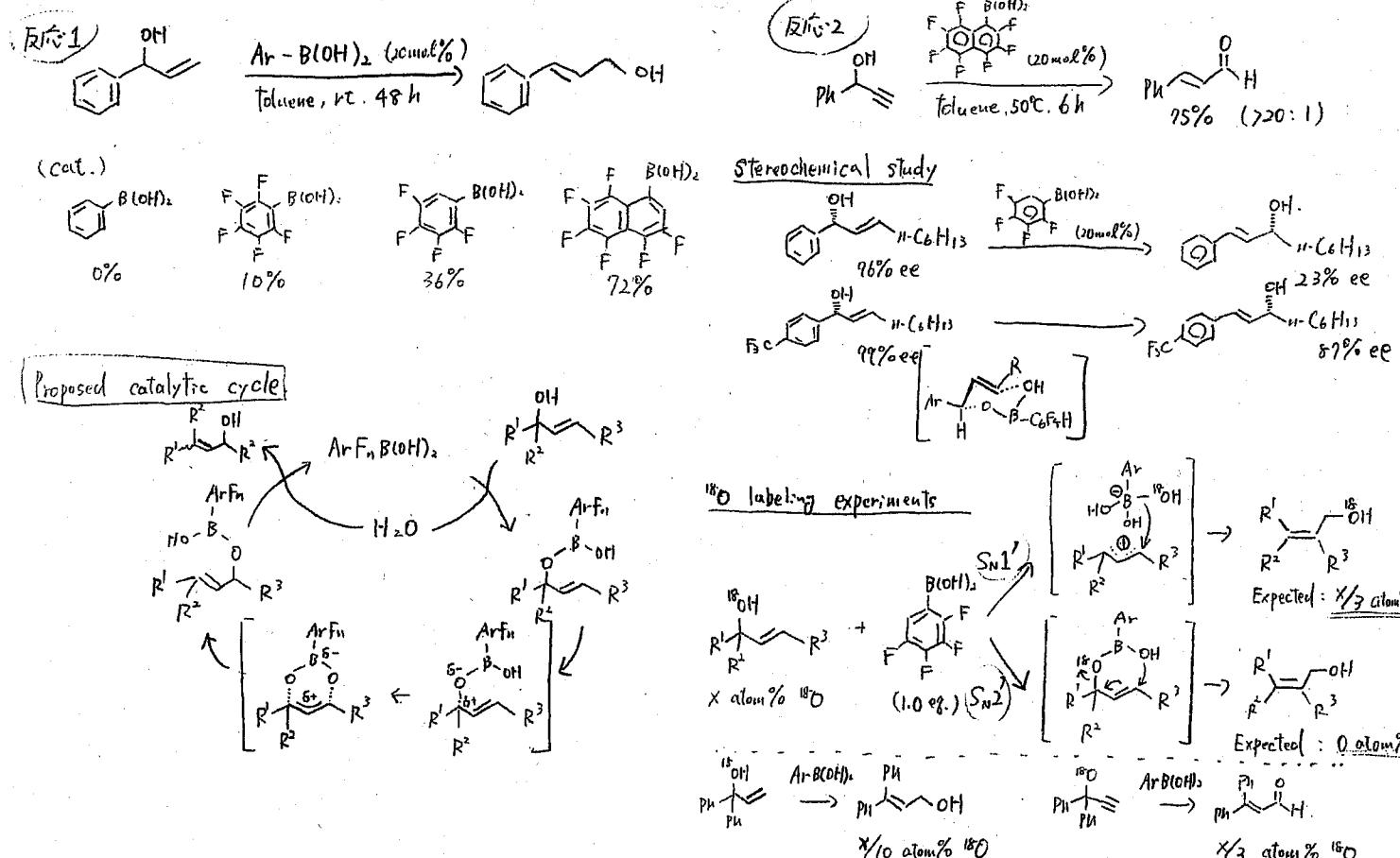


Mild and selective boronic acid catalyzed 1,3-transposition of allylic alcohols
and Meyer-Schuster rearrangement of propargylic alcohols

(Chem. Sci. 2011, 2, 1305-1310.)

Dennis G. Hall
(University of Alberta, Canada)

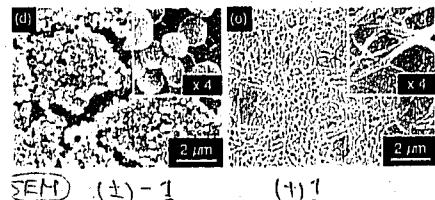
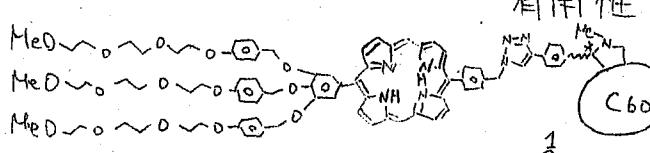


Nanorods versus Nanovesicles from Amphiphilic Dendrofullerenes

Antonio Muñoz, Beatriz M. Illescas, Macarena Sánchez-Navarro, Javier Rojo, and Nazario Martín

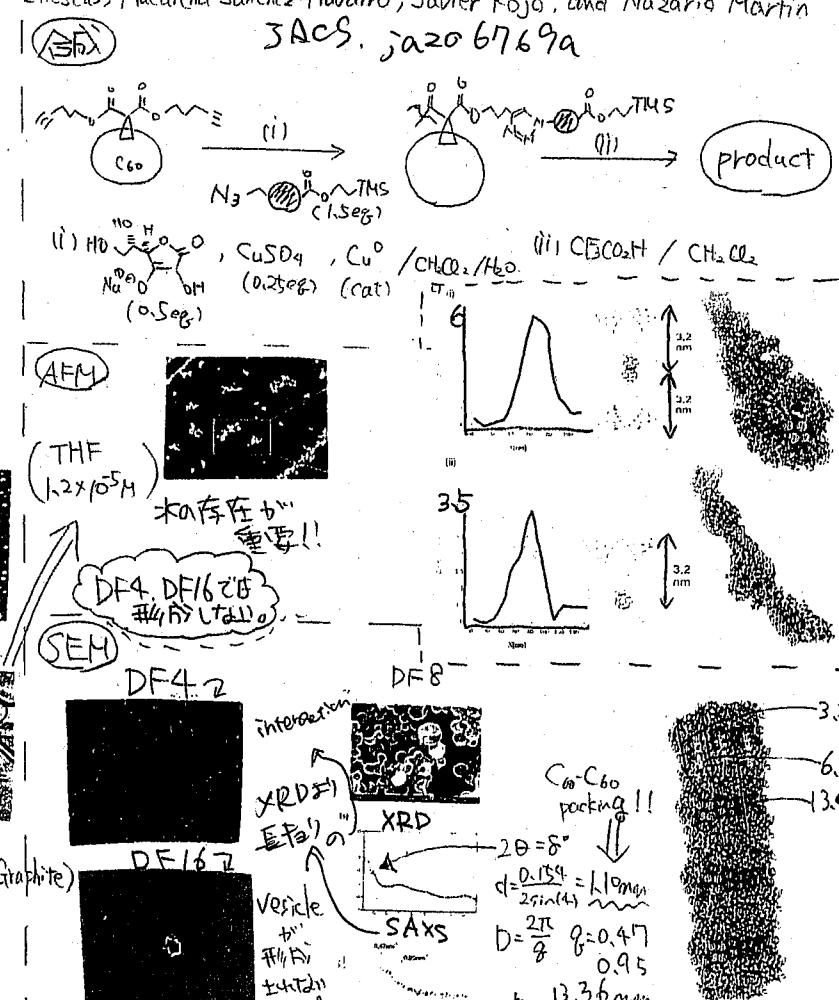
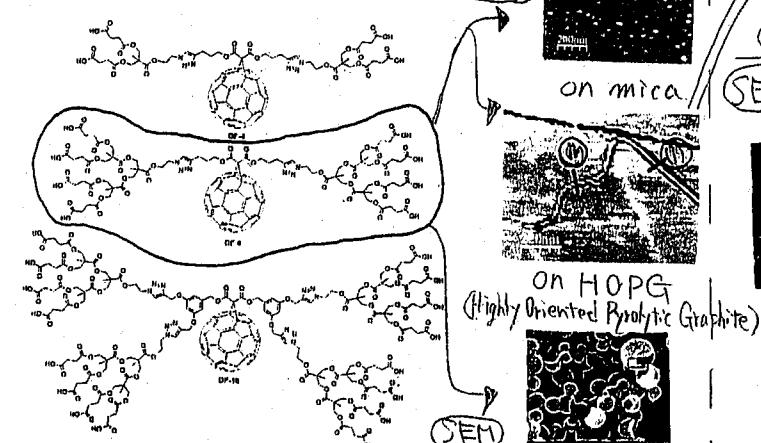
(EAT)

フラーレン誘導体の aggregation することの有用性



This work

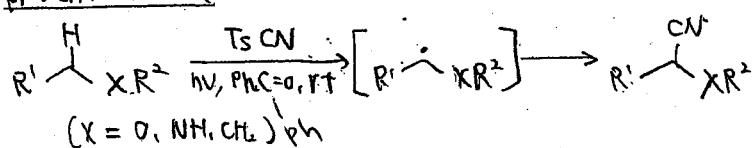
+11ボキシル基の置換C₆₀誘導体



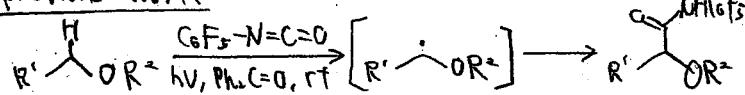
Photochemically Induced Radical Transformation of C(sp³) - H Bonds to C(sp³) - CN Bonds (D1 早)

M. Inoue, et al (Tokyo Univ.) Org. Lett. ASAP
(DOI: 10.1021/acs.orglett.5b01594)

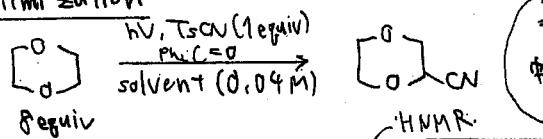
present work.



previous work.



Optimization

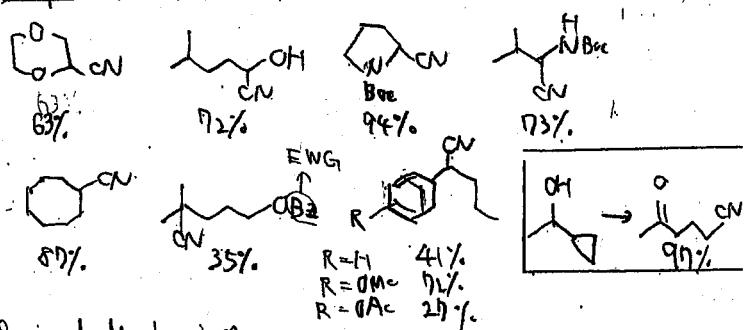


entry	solvent	Ph-C=O(equiv.)	t, (h)	yield (%)	Ph-C=O(mol%)
1	benzene	1	6	74	59
2 ^a	benzene	0	12	0	—
3	benzene	0.5	12	64	<4
4	benzene	1	24	46	48
5	dioxane	1	1	85	90
6	MeCN	1	6	72 (63)	87
7	MeCN	0.5	6	74 ↑	88

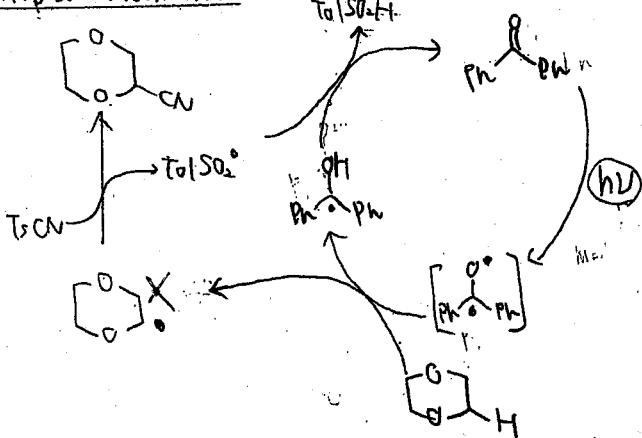
(a) TsCN (83%) D4R.

相対活性の下の単離率の値

Scope (substrate: TsCN : PhC=O = 8 : 1 : 1) in MeCN



Proposed Mechanism

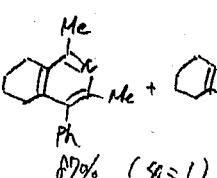
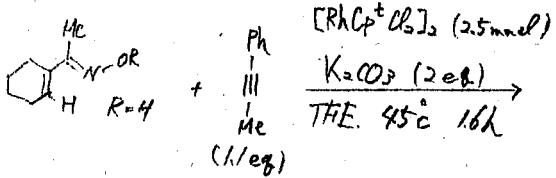


Pyridine synthesis from oximes and alkynes via rhodium(III) catalysis

:Cp* and Cp' provide complementary selectivity

T. K. Hyster, T. Rovis

M2. 調査



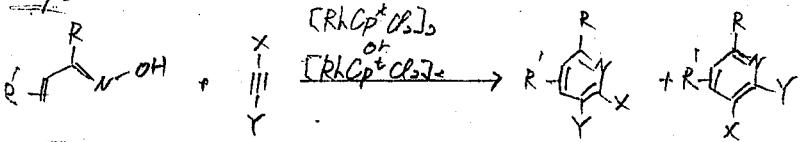
Chem. Commun. DOI: 10.1039/c1cc15248c

R = Me & Ph no reaction

Cp* と R-C≡N = Cp* と 他の生成比が逆転
83% (1:2)



Scope



Oxime

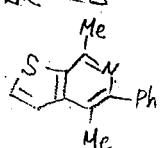
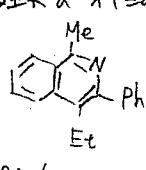
RH 主に Phキル, H₂反応が進行 (収率低)

R' が アルキル, Z=IL → 選択性低

R' が 未引基 (CH₃, IRPIL) と IL が反応しない

置換基 R' の選択性高 (6.6~10:1)

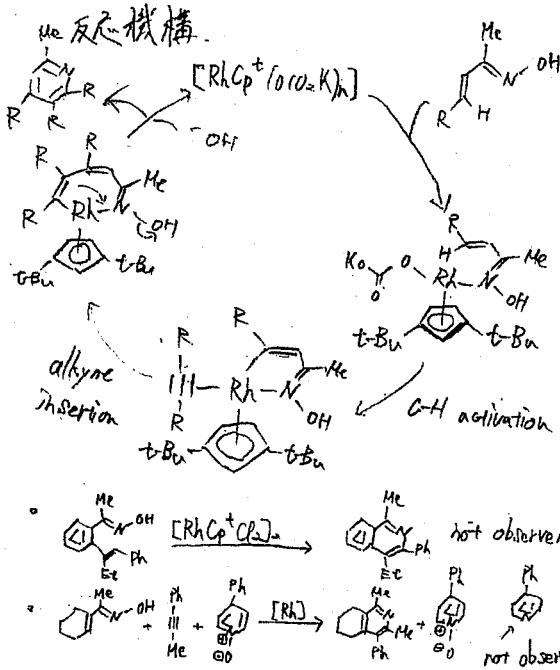
芳香環 α, β, γ 位で反応進行



alkyne $\approx \text{Cp}^*\text{I}^+ (\text{Z}=\text{Ph}, \text{Z}=\text{C}_6\text{H}_5)$
立体、電子的の差と分子
選択性

ex X = Ph, Y = C-Pt (R, R = 1)
= Ph = CO₂Me (5 = 1)
= C₆H₅CO₂ = Me (3.5 = 1)
= Me = i-Pr (10 = 1)

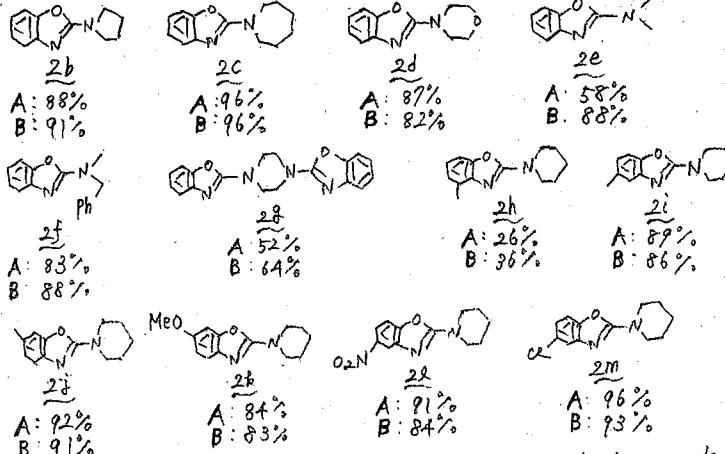
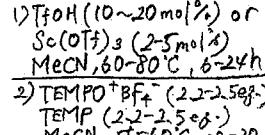
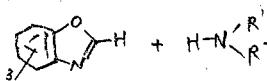
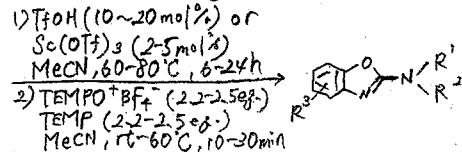
全体的に Cp* を使った方が
生成比は有利なことが多い



Organic Oxidant

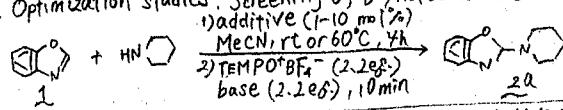
Sebastian Wertz, Shintaro Kodama, and Armido Studer*

Angew. Chem. Int. Ed (DOI: 10.1002/anie.201104735)



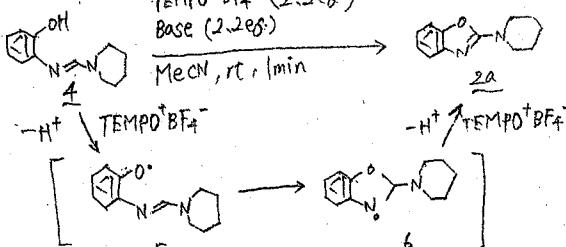
Scheme 1. Strategies for and key intermediates in the direct amination of benzoxazoles.

Table 1. Optimization studies: Screening of Brønsted and Lewis acids.



entry	Additive (mol %)	T (°C)	Base	Yield (%)
1	2,6-dCl ₂ C ₆ H ₃ CO ₂ H (10)	rt	2,6-lutidine	72
2	2,6-dCl ₂ C ₆ H ₃ CO ₂ H (10)	rt	TEMP	76
3	2,6-dCl ₂ C ₆ H ₃ CO ₂ H (10)	60	TFMPP	83
4	TfOH (10)	60	TEMP	93
5	Sc(OTf) ₃ (2)	60	TEMP	90

Scheme 2. Substrate scope: Variation of the amine and the benzoxazole core



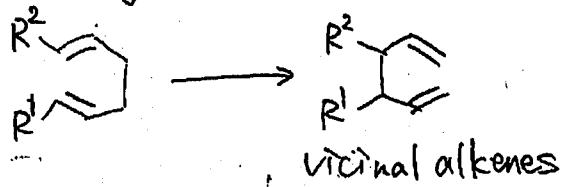
Scheme 3. Mechanistic studies

Diastereocontrol in Asymmetric Allyl-Allyl Cross-Coupling: Stereocontrolled Reaction of Prochiral Allylboronates with Prochiral Allyl Chlorides

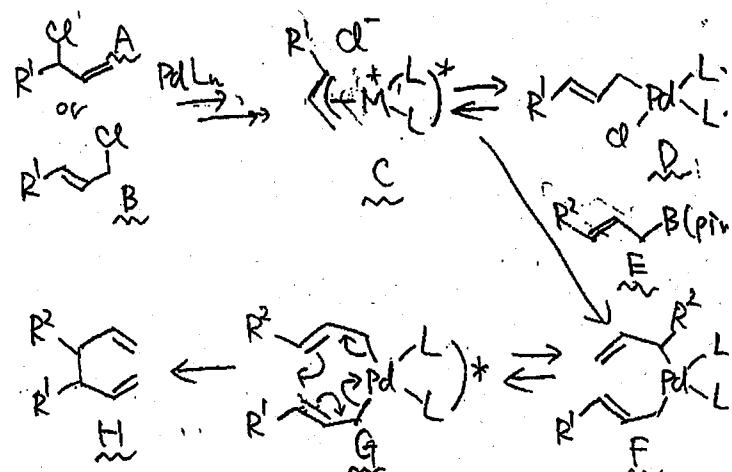
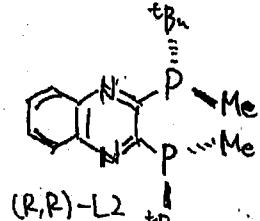
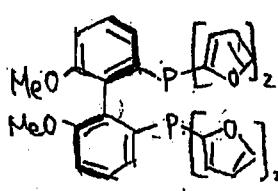
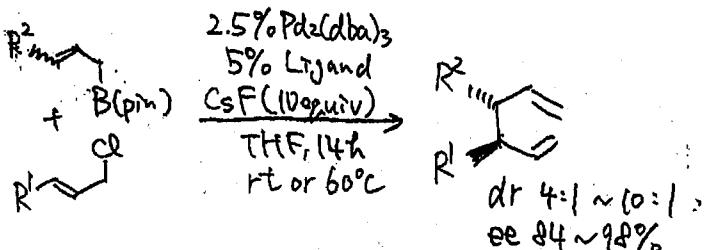
Brozek, J. A.; Ardolino, M. J.; Morken, J. P.

JACS doi: 10.1021/ja2075967

Cope rearrangement.



no chantis selective viridinf



X-ray Structure

of $\text{Li}\cdot\text{PdCl}_2$

PdE 中心上傳

图形から高さを求める



Efficient Ternary Blend Bulk Heterojunction Solar Cells with Tunable Open-Circuit Voltage

Thompson, B. et al. J. Am. Chem. Soc. 2011, 133, 14534-14537

binary blend: one donor + one acceptor

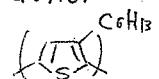
ternary blend: 1d + 2a or 2d + 1a

binary blend の太陽電池は 10 ~ 12% の変換

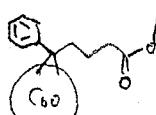
効率が限界と言われている。

This work

donor



Two acceptor

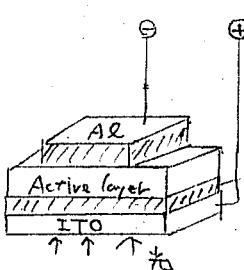


P3HT
LUMO 3.2 eV
Eg 5.2 eV
HOMO 5.9 eV

PCBM
LUMO 4.0 eV
Eg 4.2 eV
HOMO 5.9 eV

ICBA
LUMO 4.2 eV
Eg 5.9 eV
HOMO 6.0 eV

device



J_{SC} : 短絡電流密度

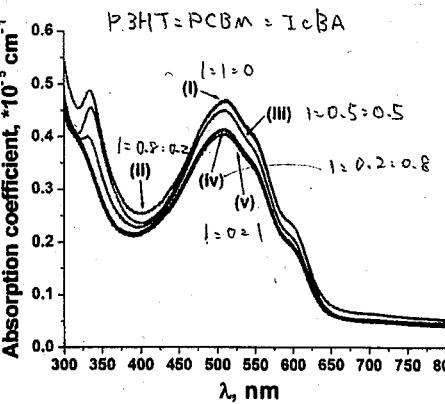
V_{OC} : 開放電圧

FF : 曲率因子

PEDOT:PSS

η : 光電変換効率

P3HT:PC ₆₁ BM:ICBA	J_{SC} (mA/cm ²)	V_{OC} (V) ^b	FF	η (%)	M1 河合
1:1:0 ^c	9.90	0.605	0.60	3.57	J_{SC} 大
1:0.9:0.1 ^d	9.22	0.618	0.59	3.29	V_{OC} 小
1:0.8:0.2 ^d	9.11	0.631	0.57	3.28	
1:0.7:0.3 ^d	8.58	0.649	0.58	3.22	
1:0.6:0.4 ^d	8.31	0.669	0.58	3.11	
1:0.5:0.5 ^d	8.27	0.688	0.57	3.18	
1:0.4:0.6 ^d	8.18	0.709	0.57	3.22	
1:0.3:0.7 ^d	8.14	0.741	0.57	3.34	
1:0.2:0.8 ^d	8.19	0.769	0.59	3.69	
1:0.1:0.9 ^d	8.18	0.804	0.60	3.91	
1:0:1 ^d	8.23	0.844	0.58	3.98	



EQE = 外部量子効率
吸収した光子が生れた電子の割合

ICBA 増 → EQE 低

UV-vis (可視光領域)

ICBA 増 → 吸光率低

光子によって生成する電子が減少したため J_{SC} が下がる。

筆者らによる仮説

Ternary system: D₁, D₂ + A (fullerene)

D₁ + A = $V_{OC} = 0.8V$ $J_{SC} = 8.0mA/cm^2$, FF=0.6, $\eta = 3.8\%$
D₂ + A = " 0.5V " 14.0mA/cm² FF=0.6 $\eta = 4.2\%$

D₁ + D₂ + A = " 0.65V " 15mA/cm² FF=0.6 $\eta = 5.9\%$

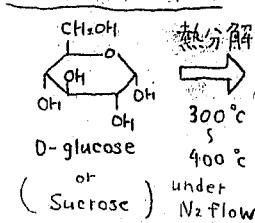
Facile synthesis of a sulfonated carbon-silica-meso composite and mesoporous silica

Devaki Nandan, Peta Sreethivasulu, Sandeep K. Saxena and Nagabhushna Viswanadham*

Chem. Commun., 2011, 47, 11537

M1 分析

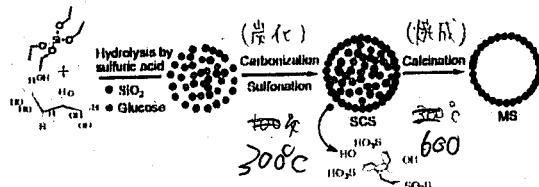
Sugar Catalyst



* 热的、化学的な安定性、强い固体酸、リサイクル可能

Toda et al. Nature, 2005, 438, 178

Present Work



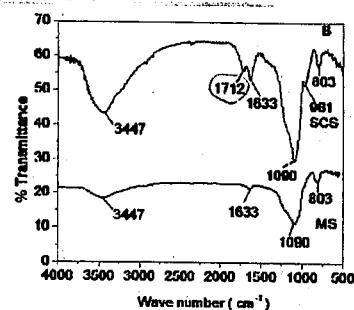
TEOS に対して 0.385 glucose : 4.8 H₂O : 0.88 H₂SO₄ (mole eq)

界面活性剤を用いてメタロ-ラスミリカを合成

→ XRD, N₂ 吸脱着等温线、SEM によると孔を確認

Sample	比表面積/m ² ·g ⁻¹ (×)	ミクロ	比体積/m ³ ·g ⁻¹	細孔径/nm
SCS	539.23	240.44	0.5 ²	7.6
MS	656.47	0	0.87	5.3

IR



SCS

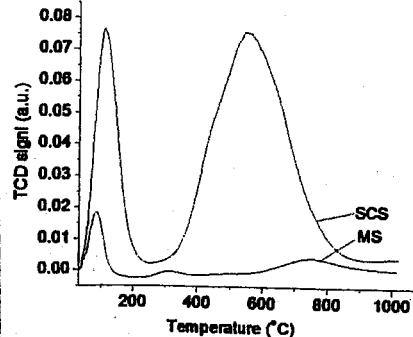
-O-SO₃H : 1712 cm⁻¹

* 1207 cm⁻¹ は -O-SO₃H 由来のπ-π*吸収であるが、三元由来のπ-π*を被って見えない。

MS

-O-SO₃H 基は消失。

TPD → アニモニアの脱離量・温度を見た。



高温で脱離 → 強い酸点

低温で脱離 → 弱い酸点

SCS 強い酸点をもつ。

MS 弱い酸点をもつ。

Phenol + t-BuOH $\xrightarrow[120^\circ C, 9 h]{SCS}$ TBP (tert-butyl phenol)

cycle	conv. of phenol (%)	Selectivity of Product (%)		
		2-TBP	4-TBP	2,4-DTPB
1	50	52	30	18
2	48	52	31	17
3	49	53	29	18

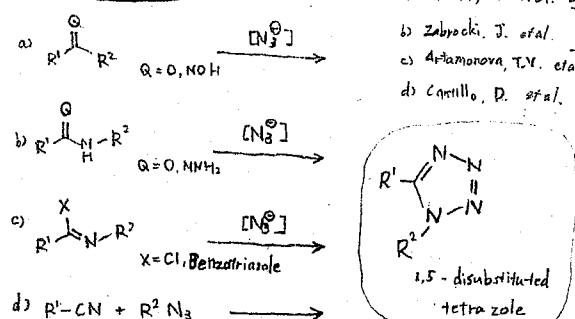
Implanting Nitrogen into Hydrocarbon Molecules through C-H and C-C Bond Cleavages:

A Direct Approach to Tetrazoles

Feng Chen, Chiang Qin, Yuxin Gui, and Ning Jiao*
Angew. Chem. Int. Ed., 2011, 50, 1. DOI: 10.1002/anie.201105505

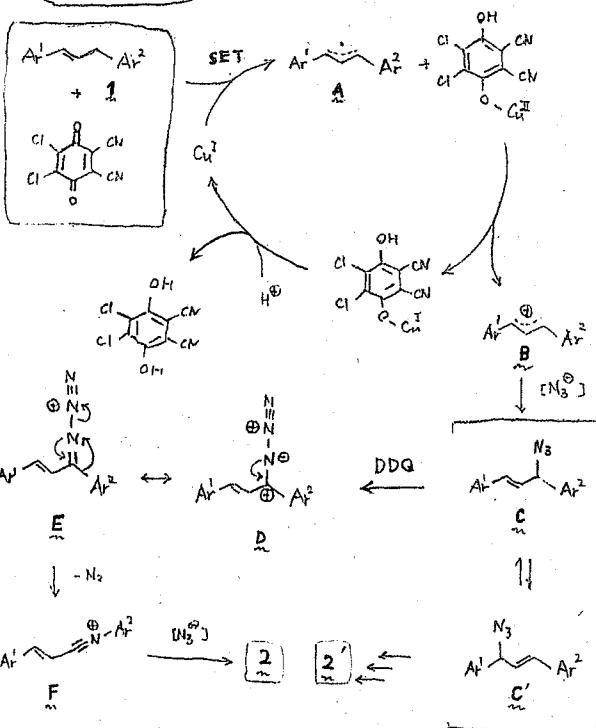
MI 矢野

テトラゾールの合成法

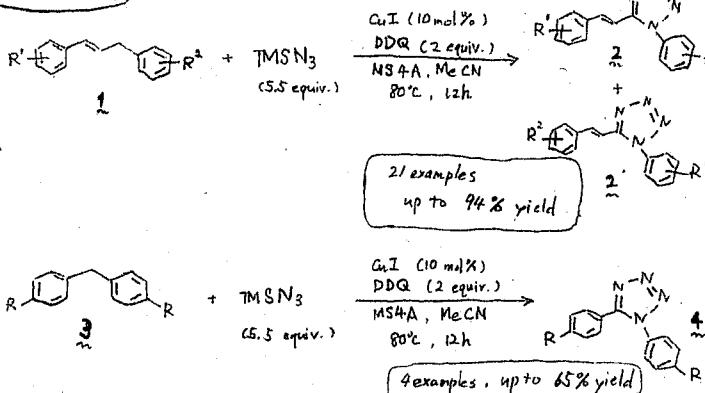


- a) Suzuki, H. et al. *Synthesis* (1993)
 b) Zabrocki, J. et al. *J. Org. Chem.* (1992)
 c) Afanorova, T.V. et al. *Synthesis* (1996)
 d) Camillio, D. et al. *J. Am. Chem. Soc.* (2011)

Proposed Pathway



本論文の手法

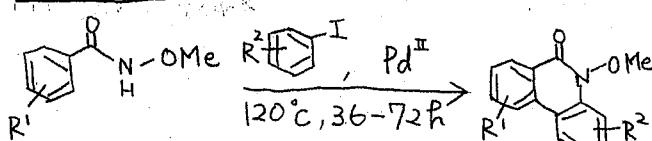


34

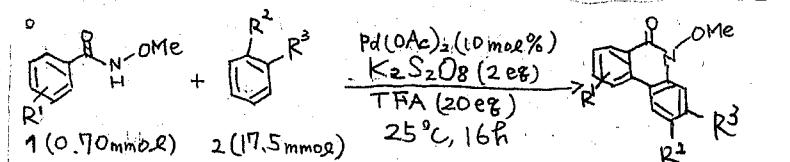
Synthesis of Phenanthridinones from N-Methoxybenzamides and Arenes by Multiple Palladium-Catalyzed C-H Activation Steps at Room Temperature.

- 2a: $\text{R}^2 = \text{H}, \text{R}^3 = \text{CH}_3$
 2b: $\text{R}^2 = \text{H}, \text{R}^3 = \text{OEt}$
 2c: $\text{R}^2 = \text{CH}_3, \text{R}^3 = \text{CH}_3$

Previous work



- expensive metal oxidant
- high temperature
- long reaction time



1	2	Yield [%]	1	2	Yield [%]
		92			64
		80			86
		90			74
		86			86
		89			83

